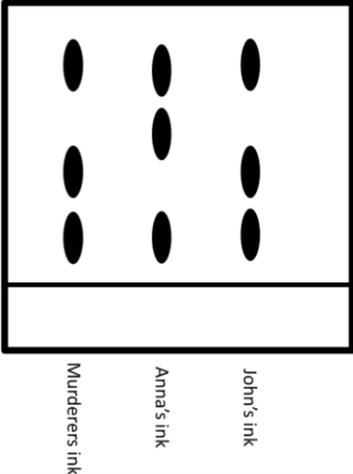
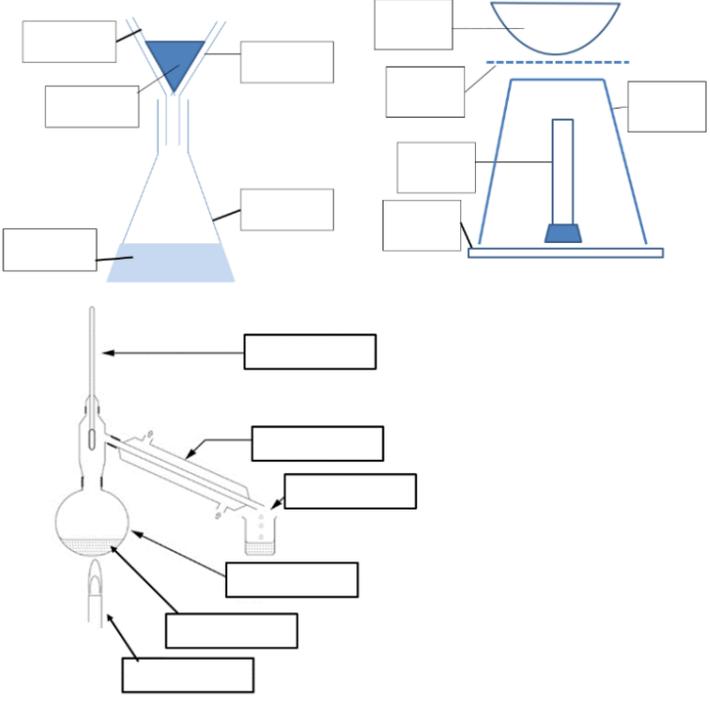
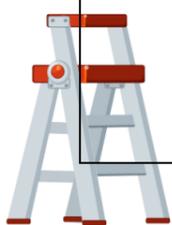


Percentage	I can ...	Prove it!
	<p>Link neutralisation to concentration of contributing acids and alkalis.</p> <p>Use knowledge of solubility to separate salt from muddy sea water</p> <p>Analyse the colours in an unknown sample from a chromatograph.</p> <p>Write a symbol equation for neutralisation.</p>	<ol style="list-style-type: none"> 1) Define concentration and explain why it is important to know the concentration of an acid when trying to neutralise an alkali. 2) Write a method for how we could get pure salt left over from a mixture of mud and sea water. 3) Decide who committed the murder based on the chromatograph. Justify your answer. 4) Write a symbol equation for the neutralisation of hydrochloric acid and sodium hydroxide. 
	<p>Explain how acid rain could be reduced.</p> <p>Write a word equation for neutralisation</p> <p>Explain how a condensing tube makes distillation more effective</p> <p>Describe the effect of heat on solubility</p>	<ol style="list-style-type: none"> 1) Explain in detail 3 steps that could be taken to reduce acid rain. 2) Write a word equation for the neutralisation of hydrochloric acid with sodium hydroxide. 3) Explain in terms of particles and temperature what a condensing tube does to the gas passing through it. 4) Describe the effect of heat on the rate at which something dissolves.
	<p>Name acids and alkalis from their formulae</p> <p>Link acid rain to combustion of fossil fuels.</p> <p>Write a method for neutralisation.</p> <p>Explain how filtration, evaporation and distillation work using the correct key terms</p>	<ol style="list-style-type: none"> 1) Give the names for the following acids and bases from their formulae <ol style="list-style-type: none"> a. HCl b. H₂SO₄ c. NaOH 2) Write a note to Mr Holmes explaining in 3 steps how combustion of fossil fuels causes acid rain. 3) Mr Trueman has got stomach acid and it is causing him a lot of pain. Explain to him what he could do to make himself feel better making reference to acids and alkalis. 4) Label these diagrams with the equipment needed and explain how each separation technique works using the following key terms. <i>Solute, residue, filtrate, solvent, solution, mixture, evaporation, condensation, filtration, insoluble, soluble</i> 



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	<p>Explain how the pH scale can be used to determine if a liquid is an acid or a base.</p> <p>Explain how litmus paper can be used to test for acids and bases.</p> <p>Describe how neutralisation is achieved</p> <p>Decide which separation method is the best to use in different scenarios.</p>	<p>1) You test the pH of three liquids, state whether they are an acid, a base, or a neutral liquid.</p> <p>a) pH=2 b) pH=8 c) pH=7 d) pH=10</p> <p>2) When blue litmus paper is added to an acid it turns _____. When red litmus paper is added to an alkali it turns _____.</p> <p>3) Mr Findlay is trying to neutralise an acid. Write him a note describing how he could do it and what he might have to use.</p> <p>4) Write the preferred separation method next to the substance.</p> <table border="1" style="width: 100%;"> <tr> <td style="width: 70%;"></td> <td>Two liquids with different boiling points</td> </tr> <tr> <td></td> <td>An insoluble solid from a liquid</td> </tr> <tr> <td></td> <td>A soluble solid from a liquid</td> </tr> <tr> <td></td> <td>Two or more liquids with different colours</td> </tr> </table>		Two liquids with different boiling points		An insoluble solid from a liquid		A soluble solid from a liquid		Two or more liquids with different colours																	
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	<p>Describe how to use universal indicator and the results that you may expect to get.</p> <p>Identify one effect of acid rain</p> <p>Define neutralisation.</p> <p>Define key words to do with solubility</p> <p>Identify the equipment needed for each separation technique.</p>	<p>1) Label the pH scale with acid alkali and neutral</p> <p>2) What is one effect of acid rain?</p> <p>3) When an acid is added to an alkali, we call this _____</p> <p>4) Link the key words to their meanings</p> <table border="1" style="width: 100%;"> <tr> <td style="width: 50%;">Solute</td> <td style="width: 50%;">The liquid that does the dissolving</td> </tr> <tr> <td>Solvent</td> <td>The soluble solid</td> </tr> <tr> <td>Soluble</td> <td>A special type of mixture made when a solid dissolves and mixes with a liquid</td> </tr> <tr> <td>Insoluble</td> <td>A solid that <u>can</u> dissolve</td> </tr> <tr> <td>Solution</td> <td>A solid that <u>cannot</u> dissolve</td> </tr> </table> <p>5) Choose from the equipment list below the relevant equipment needed to do each of the separation techniques.</p> <p>a. Filtration b. Distillation c. Chromatography d. Evaporation</p> <table border="1" style="width: 100%;"> <tr> <td>Heat proof mat</td> <td>Tripod</td> <td>Bunsen burner</td> <td>Gauze</td> <td>Conical flask</td> </tr> <tr> <td>Evaporating basin</td> <td>Condensing tube</td> <td>Beaker</td> <td>Filter paper</td> <td>Funnel</td> </tr> <tr> <td>Round bottom flask</td> <td>thermometer</td> <td>Balloon</td> <td>Tongs</td> <td>Spatula</td> </tr> </table>	Solute	The liquid that does the dissolving	Solvent	The soluble solid	Soluble	A special type of mixture made when a solid dissolves and mixes with a liquid	Insoluble	A solid that <u>can</u> dissolve	Solution	A solid that <u>cannot</u> dissolve	Heat proof mat	Tripod	Bunsen burner	Gauze	Conical flask	Evaporating basin	Condensing tube	Beaker	Filter paper	Funnel	Round bottom flask	thermometer	Balloon	Tongs	Spatula
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Key Terms:

Chemical Reversible Solute Solvent soluble insoluble Solution Dissolve Mixture Separate Acid
 alkali neutralisation acid rain fossil fuel pollution litmus paper universal indicator pH scale
 Distillation Evaporation Filtration funnel filtrate residue

